	Sanjay Ghodawat University, Kolhapur Established as State Private University under Govt. of Maharashtra. Act No XL, 2017	2023-24 EXM/P/09/00

Year and Program: **First Year (B. Pharm)-2023-24**	**School of Pharmaceutical Sciences**	**Department : Pharmacy**
Course Code : BP 104T	**Course Title :Pharmaceutical Inorganic Chemistry I**	**Semester – I**
Date: 18/12/2023 **Day :** Monday **Time:** 10.30 am to 1.30 PM	**Examination: End Semester**	**Max Marks: 75**

Instructions: 1) All questions are compulsory.
2) Draw a neat labelled diagram if necessary.

		Marks	Bloom's Level	CLO
Q.1	**Objective Type Questions (20 x 1)**	**20**		
a.	Which is both acidifying as well as expectorant: a. Ammonium chloride b. Potassium iodide c. Dil. HCl d. Sodium potassium Tartrate	01	01	CLO3
b.	According to Arrhenius theory base is a substance which gives…… when dissolved in water a. Hydroxyl ion b. Hydronium ion c. Hydrogen ion d. None of these	01	01	CLO2
c.	Boric acid is used as: a. Urinary acidifier b. Antimicrobials c. Antioxidants d. analgesics	01	01	CLO3
d.	The category of calcium gluconate is: a. antacid b. calcium replenisher c. antioxidant d. radiopharmaceuticals	01	01	CLO3
e.	The concentration of standard iron solution in the limit test for iron is: a. 10 ppm b. 15 ppm c. 20 ppm d. 25 ppm	01	01	CLO1

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f.	Limit tests are performed in: a. Nessler's cylinder b. Measuring cylinder c. Volumetric Flasks d. Iodine flasks	01	01	CLO1
g.	Limit tests are …… test designed to identify small quantities of impurities a. Qualitative b. Quantitative or semi quantitative c. Imperative d. All of the above	01	01	CLO1
h.	According to Lewis theory, acid is… a. Electron pair acceptor b. Electron pair donor c. Proton acceptor d. Electron acceptor	01	01	CLO2
i.	Boron trifluoride is… a. Acid b. Base c. Neutral d. amphoteric	01	01	CLO2
j.	The p^H of blood is… a. 7.4 to 7.5 b. 5.4 to 7.5 c. 4.5 to 8.0 d. 2.0 to 4.0	01	01	CLO2
k.	Carbonic acid is ……. Acid. a. Strong b. Weak c. Very strong d. None of these	01	01	CLO2
l.	Boric acid is having …… mol. Weight. a. 61.83 b. 40.10 c. 45 d. 50	01	01	CLO3
m.	Backing soda is known as a. Na2Co3 b. NaHCo3 c. NaOH d. K2CO3	01	01	CLO3
n.	Hydrogen peroxide is used as… a. Antiseptic b. Acidifying agent	01	01	CLO3

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c. Protective
d. antioxidant

o.	Which of the following is expectorant drug? a. Copper sulphate b. Potassium iodide c. Ferrous sulphate d. Zinc sulphate	01	01	CLO
p.	Dentifrices are used for… a. Antacid b. Emetics c. Dental product d. Antimicrobial	01	01	CLO
q.	..... is/are sources of impurities a. Raw material b. Method of manufacturing c. Both a and b d. None of these	01	01	CLO
r.	The first edition of IP got published in… a. 1956 b. 1955 c. 1958 d. 1947	01	01	CLO
s.	The strength of a Bronsted acid depends upon its tendency to ..... a proton a. Gain b. Accept c. Both a. and b. d. None of the above	01	01	CLO
t.	One of the following limit tests is based on the comparison of colour a. Limit test for sulphate b. Limit test for chloride c. Limit test for iron d. Both a and c	01	01	CLO1

Q.2	**Long Answers (Answer any 2 out of 3)**	**20**		
a.	Write a note on history of Indian Pharmacopoeia. Highlight the salient features of I.P.2018.	10	01	CLO1
b.	Write principle & reaction involved in a) Limit test for Arsenic b) Limit test for Iron	10	01	CLO2
c.	Write a note on dental products, classify in detail.	10	02	CLO3

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Q.3		**Short Answers (Answer 7 out of 9)**	**35**		CLO
	a.	What are radiopharmaceuticals? Enumerate the properties of alpha, beta and gamma radiations	5	01	CLO4
	b.	What are cathartics. Classify and explain with suitable examples.	5	02	CLO3
	c.	Define & classify antimicrobial agents.	5	02	CLO3
	d.	Discuss the role of Fluoride in detail.	5	01	CLO3
	e.	Define tonicity. Explain Handerson-Hasselbach equation.	5	01	CLO2
	f.	Define Haematinics. Give method of preparation, properties & uses of ferrous sulphate.	5	02	CLO3
	g.	Define a) Measurement of radioactivity b) Half-Life	5	01	CLO4
	h.	Define Expectorant. Write method of preparation, properties, & uses of potassium iodide.	5	02	CLO3
	i.	Define emetics. Give method of preparation, properties& uses of copper sulphate.	5	02	CLO3

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